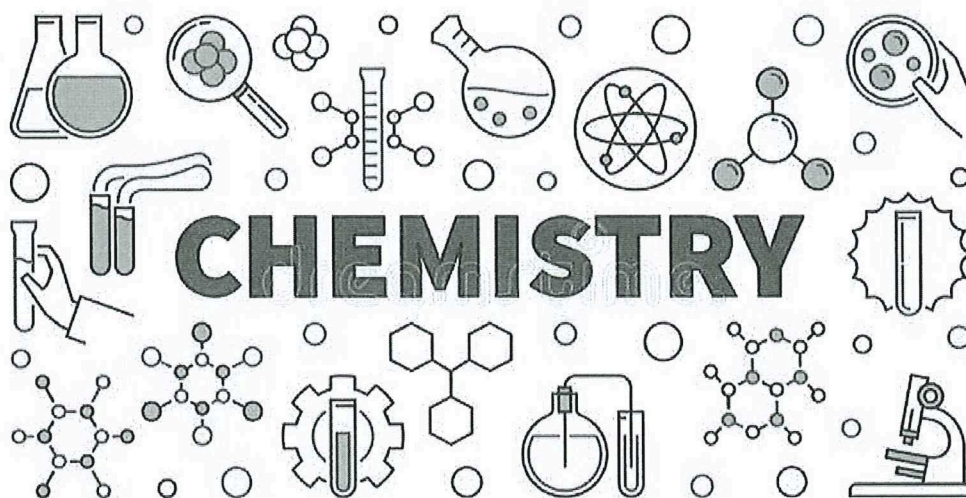


WJEC Chemistry 1
Dual Award - Higher Tier
Mark Scheme



WJEC Chemistry 1
Dual Award – Higher Tier
1.1 Mark Scheme

Question	Marking details	Marks Available					
		AO1	AO2	AO3	Total	Maths	Prac
6 (a)	<p>0.992g / 0.99g (3)</p> <p>if answer incorrect allow credit for correct working</p> <p>4 (23) Na / 92g Na (1)</p> <p>$\Rightarrow 2$ (62) Na₂O / 124g Na₂O (1)</p> <p>allow ecf</p> <p>alternative method</p> <p>0.032 mol Na (1)</p> <p>0.016 mol Na₂O (1)</p> <p>allow ecf</p>		3		3		3
(b)	<p>(i)</p> <p>34.07% / 34.1% / 34%</p> <p>allow ecf from part (a)</p> <p>(ii)</p> <p>award (1) for any of following</p> <ul style="list-style-type: none"> • sodium was tarnished / oxidised already / some of the sodium had already reacted • insufficient burning time / not all reacted • not enough oxygen used • other reactions (or named) had taken place <p>neutral</p> <ul style="list-style-type: none"> • some product was spilled • incorrect measurements 		1		1		1
	Question 6 total	0	4	1	5	4	5

Question	Marking details	Marks available					
		AO1	AO2	AO3	Total	Maths Prac	
5	(a) mixture is heated until one liquid boils / evaporates then condenses and is collected in a different container (1) liquids must have different boiling points (1) liquid E has the lower boiling point so is the one removed / liquid F has the higher boiling point so is the one left in the flask (1)	3			3		3
	(b) 74 / 74.1 (2) if answer incorrect award (1) for $\frac{32}{43.2} \times 100$		2		2	2	
	Question 5 total	3	2	0	5	2	3

HIGHER TIER ONLY QUESTIONS

Question	Marking details	Marks available					
		AO1	AO2	AO3	Total	Maths	Prac
3	(a)						
	(b)						
	(c)						
Question 3 total		0	0	4	4	0	4